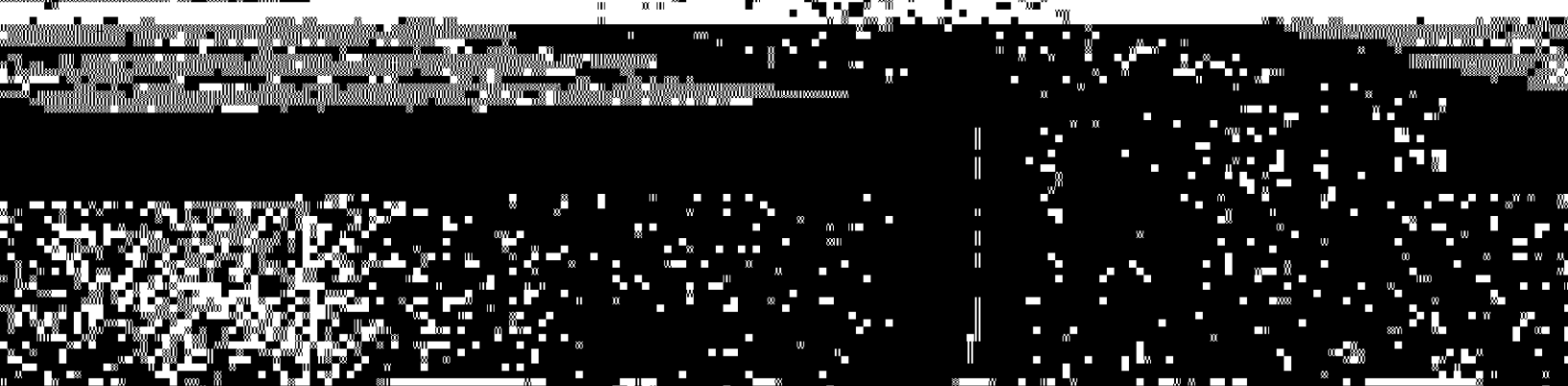
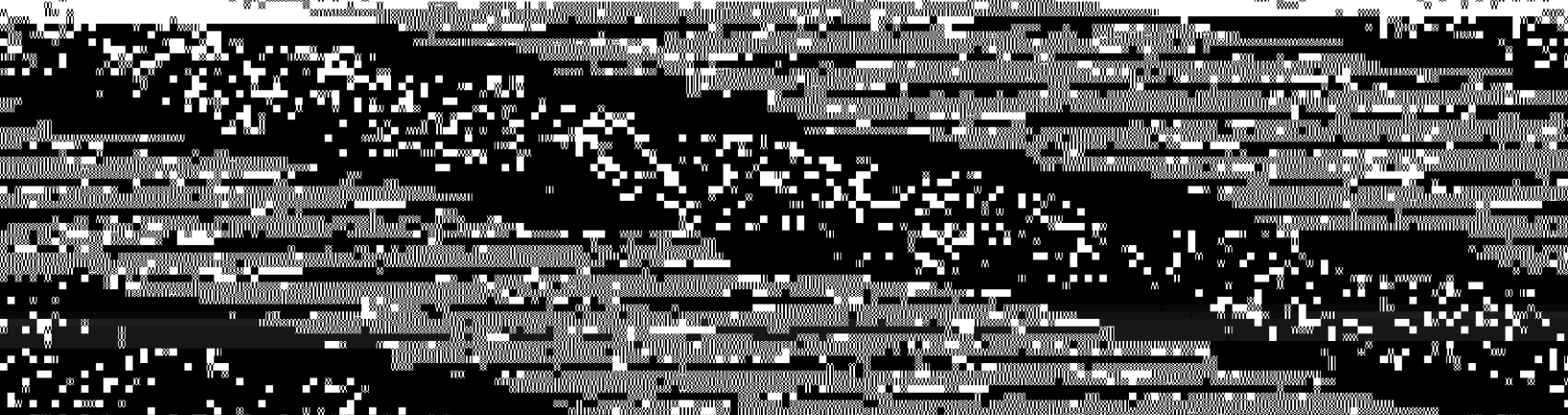


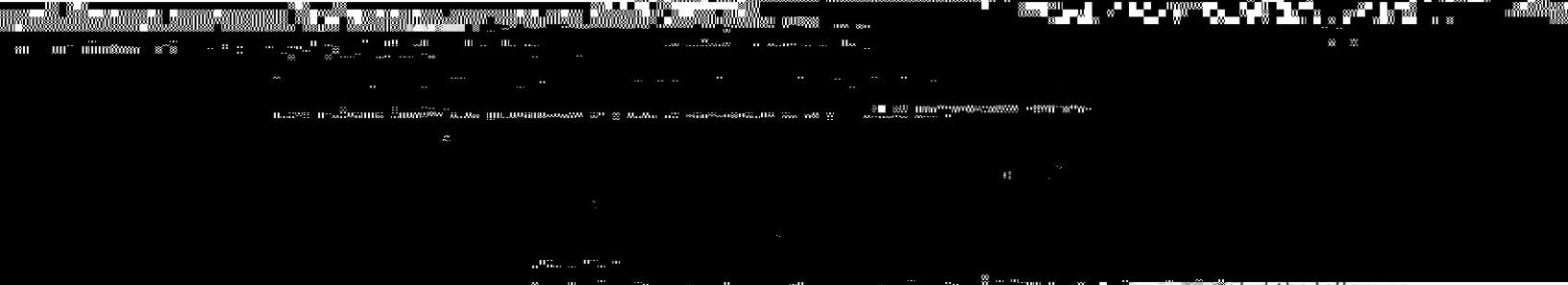
local space parts that have indicated the large degree of variability exhibited, offered by a nonlinearly optimized (exact) numerical algorithm.





...approach to the ...
 ...of large ...
 ...a direct ...
 ...to predict ...

...pressing forces that are ...
 ...in ...

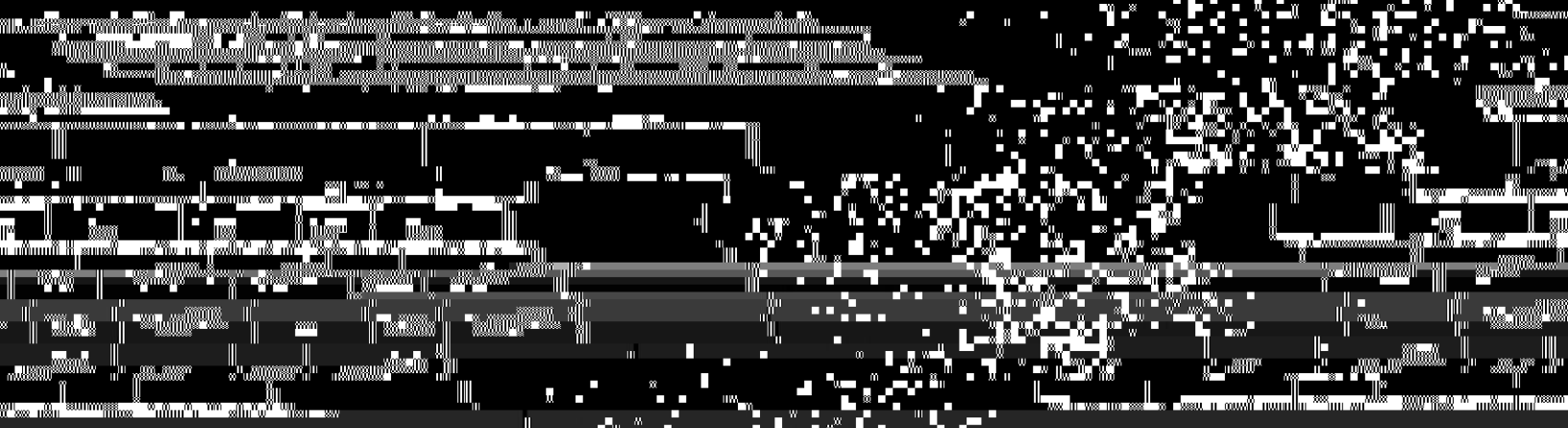


Analysis of the ... systems ... was ... the following features:

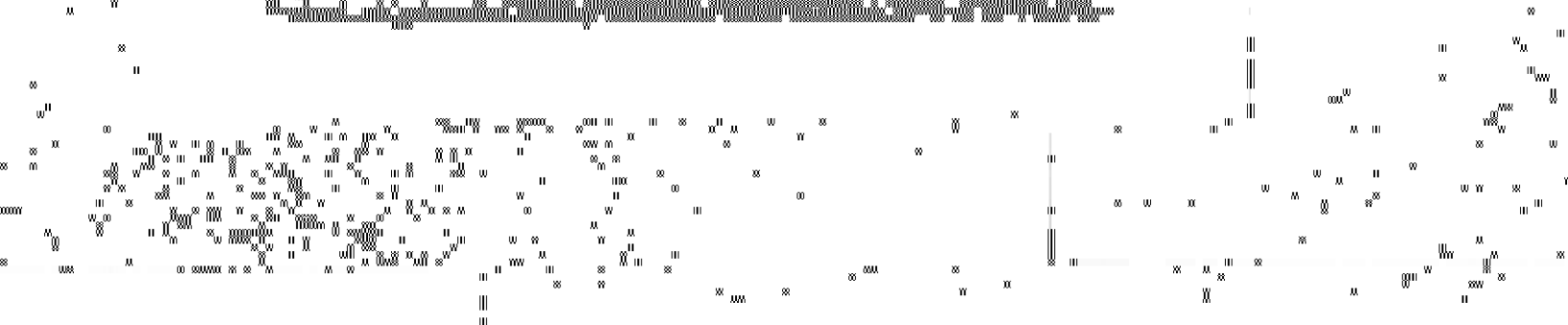
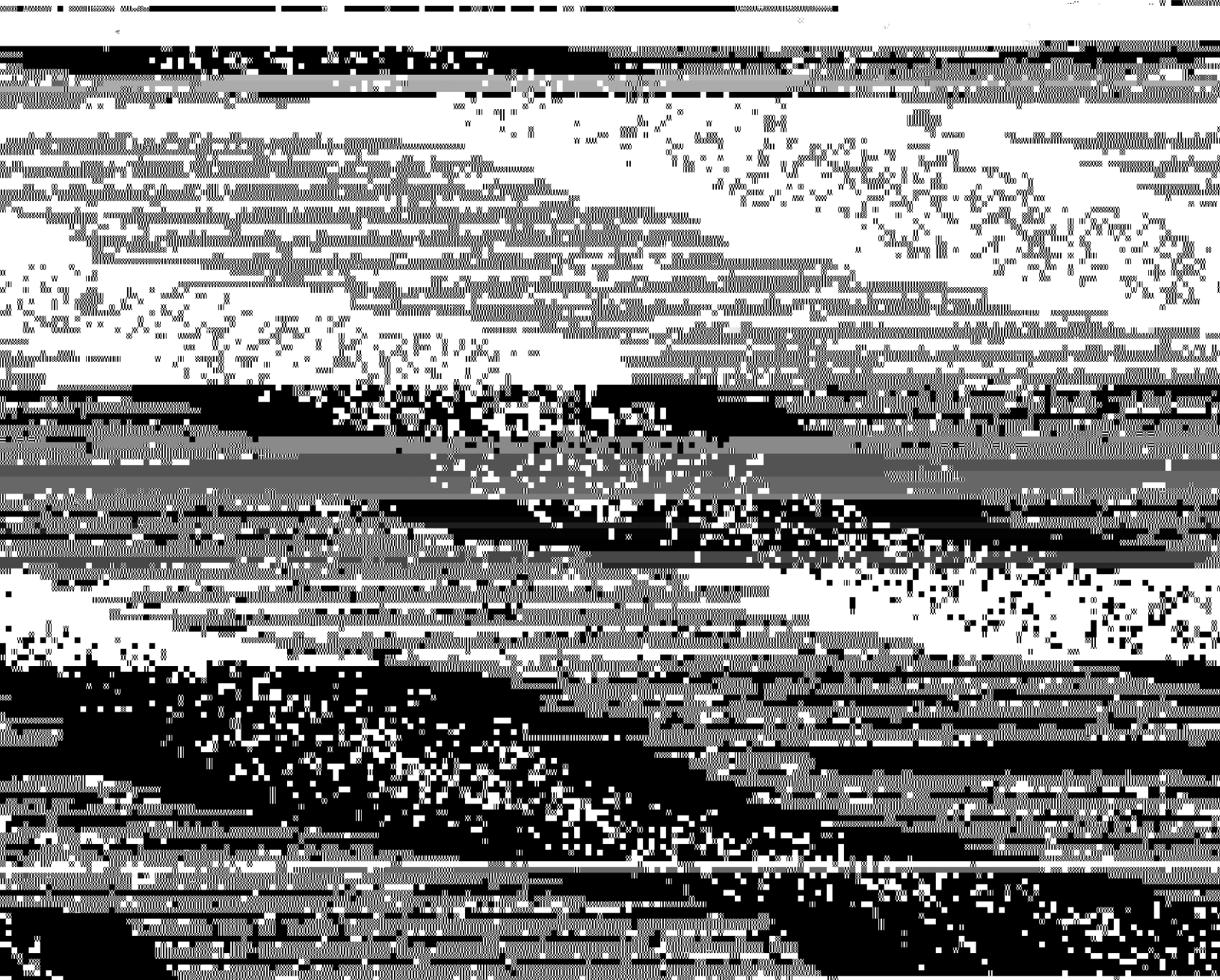
- (1) ...



... However, only an ...



the reasonable assumption of a reasonable amount of interaction energy in the polymer chains, the energy of interaction of the polymer chains in the polymer matrix.





approximate values of ρ and ρ_c are given in Table I. The values of ρ are calculated from the measured density of the polymer, and the values of ρ_c are calculated from the measured density of the polymer and the calculated crystal density.



Chemical Society Press, London, 1950) for cooling to 0°K (reference to R. Stull, Ed., *JANAF Tables of Thermochemical Data* (Dow Chemical Co., Midland, Mich., 1965)) and finally for the zero-point energy of the solid (using a Debye equation with $\theta_D = 1700^\circ\text{K}$ using data given by P. J. Gielisse, S. S. Mitchell, J. N. P. Hubbard, J. C. Van der Grinten, and R. Marshall and E. A. DiPasquale, *Phys. Rev.* **158**, 2039 (1967)). Loeb obtains an estimate of 14.4 kcal/mole for the standard enthalpy of formation at 0°K. The accuracy of this quantity is not possible to assess.

(1957) *J. Chem. Phys.* **25**, 1016 (1957)